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Solutions

Worksheet 1

**Solutions**

**Worksheet 1**

**Molarity**

**Answers**

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## **Solutions Worksheet 1 Molarity Answers**

Molarity Worksheet # 1

. 1. 15.8 g of KCl is dissolved in 225 mL of water. Calculate the molarity.  $15.8 \text{ g} \times 1 \text{ mole} / 74.6 \text{ g} = 0.212 \text{ M}$   
 $0.212 \text{ M} \times 0.225 \text{ L} = 0.0477 \text{ moles}$

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## **Molarity Worksheet**

### **# 1 Molarity Answers**

You should try to answer the questions without referring to your textbook. If you get stuck, try asking another group for help.

Calculate molarity if 25.0 mL of 1.75 M HCl diluted to 65.0 mL.

Calculate molarity by dissolving 25.0g NaOH in 325 mL of solution.

Calculate grams of solute needed to prepare 225 mL of

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0.400 M KBr solution.

Molarity Answers

**Molarity 1**

**(Worksheet) -**

**Chemistry**

**LibreTexts**

Solutions Worksheet 1

Molarity Answers 78.9

g x 1 mole. Molarity =

303.76 g = 0.519 M

0.5000 L. Solutions

Worksheet 1 Molarity

Answers Molarity

Worksheet 1 Answer

Key Chemistry Assume,

unless otherwise told,

that in all problems

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Worksheet 1

water is the solvent.

Example #1: Given a

Page 3/8.

## **Solutions Worksheet 1 Molarity Answers**

Molality Worksheet #1

Answer Key Solutions

Worksheet 1 Molarity

Answers  $78.9 \text{ g} \times 1$

mole. Molarity =

$303.76 \text{ g} = 0.519 \text{ M}$

$0.5000 \text{ L}$ . Solutions

Worksheet 1 Molarity

Answers Molarity

Worksheet 1 Answer

Key Chemistry Assume,

*Page 7/27*

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Worksheet 1

Molarity Answers

unless otherwise told,  
that in all problems  
water is the solvent.

Example #1: Given a  
Page 3/8. Solutions  
Worksheet 1 Molarity

...

## **Solutions Worksheet 1 Molarity Key**

What is the molarity of  
a solution that contains  
0.00372 moles

hydrochloric acid in  
 $2.39 \times 10^{-2}$  liters of  
solution? 0.00372 mol

HCL = 0.156 M HCL

*Page 8/27*



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$2.39 \times 10^{-2}$  L soln A  
flask contains 85.5 g  
 $C_{12}H_{22}O_{11}$  (sucrose)  
in 1.00 liters of  
solution.

## **Molarity Worksheet #1 - Science Done Wright**

Solutions What is the  
molarity of the  
following solutions  
given that: 1) 1.0  
moles of potassium  
fluoride is dissolved to  
make 0.10 L of  
solution. 1.0 mole KF =

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10. M 0.10 L soln 2) 1.0 grams of potassium fluoride is dissolved to make 0.10 L of solution.  $1.0 \text{ g KF} \times \frac{1 \text{ mole KF}}{58 \text{ g KF}} = 0.0172 \text{ mol KF}$   
 $0.0172 \text{ mol KF} / 0.10 \text{ L soln} = 0.17 \text{ M}$

**Molarity Worksheet**

**W 331 - Everett**

**Community College**

Molarity Worksheet 1

Science At Yorkdale

With Jessica Molarity

Worksheet 1 For Each

Of The Following

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Worksheet 1

Problems Use Proper  
Units And Show All  
Work 1 If 107 Grams Of  
 $\text{NH}_4\text{Cl}$  Is Dissolved In  
Enough Water To Make  
800 mL Of Solution  
What Will Be Its  
Molarity Answer 0.25  
Mol/L 2 Calculate The  
Molarity Of A.

**Molarity Worksheet  
1 Answer Key  
Chemistry**

Solution:  $MV = \text{grams} /$   
 $\text{molar mass. (x) (1.000}$   
 $\text{L)} = 245.0 \text{ g} / 98.0768$

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$\text{g mol}^{-1} \cdot x =$

2.49804235 M. to four sig figs, 2.498 M. If the volume had been specified as 1.00 L (as it often is in problems like this), the answer would have been 2.50 M, NOT 2.5 M.

### **ChemTeam: Molarity Problems #1 - 10**

A similar unit of concentration is molality (m), which is defined as the number of moles of solute per

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kilogram of solvent, not  
per liter of solution:

(15.3.1)  $m o l a l i t y =$   
 $m o l e s s o l u t e k i l$   
 $o g r a m s s o l v e n t$

Mathematical

manipulation of

molality is the same as  
with molarity.

### **15.03: Solution Concentration - Molality, Mass Percent ...**

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Molarity Answers

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Molarity Answers

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**Worksheets**

Molar Concentration of  
Solutions 1. What is the

molarity of a solution  
made by dissolving

3.00 moles of NaCl in  
enough water to make

6.00 liters of solution?

2. What is the molarity  
of KCl solution

containing 1.70 moles  
of KCl in 3.00 liters of

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### Worksheet 1

solution? 3. What is the molarity of a solution containing 4.20 moles of sulfuric acid in 300.0 mL of ...

### **Molar Concentration of Solutions**

1 mol  $\text{CaCO}_3$  100.0 g

$\text{CaCO}_3 = 0.500 \text{ mol}$

$\text{CaCO}_3? \text{ L} = 500.0 \text{ mL}$

$\times 1 \text{ L} 1000 \text{ mL} = 0.500$

$\text{L M} = 0.500 \text{ mol } 0.500$

$\text{L} = 1.00 \text{ M M} = 6.0 \text{ mol}$

$4.0 \text{ L} = 1.5 \text{ M}$  7. How

many liters of solution

can be produced from



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2.5 moles of solute if a  
2.0 M solution is  
needed?  $2.0 \text{ M} = 2.5$   
moles liters of solution  
liters of solution = 1.25  
 $L = 1.3 \text{ L}$  8.

**Molarity: Molarity =**  
**1. 2. - Central Bucks**  
**School District**

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I have two solutions. In  
the first solution, 1.0  
moles of sodium

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#### Molarity Answers

chloride is dissolved to make 1.0 liters of solution. In the second one, 1.0 moles of sodium chloride is added to 1.0 liters of water. Is the molarity of each solution the same?

### **Solutions Worksheet 1 Molarity Answer Key**

Molar concentration (also called molarity, amount concentration or substance

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Worksheet 1

Molality Answers

concentration) is a measure of the concentration of a chemical species, in particular of a solute in a solution, in terms of amount of substance per unit volume of solution Molality worksheet #1 answer key. Molality worksheet #1 answer key

## **Molality Worksheet #1 Answer Key**

Key+. 1)++23.5g+of+ NaCl+isdissolvedineno

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ughwatertomake.683L  
ofsolution. + a)+What  
+is+themolarity)(M)+o  
f+the+solution?+++ M  
olar+mass+of+NaCl+  
=58.44g/mole+  
Moles+of+NaCl:+ 23.5  
g+NaCl+++1moleNaCl  
+++ =+++ .402moles+N  
aCl+ ++++++  
+++++5  
8.44gNaCl+ ++ Molarit  
y+++ =+++++  
+moles+++++  
+++++ =+++++0.4  
02moles+NaCl+++++  
++ =0.589moles+NaCl/

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Worksheet 1

Molarity Answers  
Solution 0.683 L of solution  
+ + b) + + How + many +  
moles + of + NaCl + are co  
ntained + in + 0.0100 + L o  
f + the + above + NaCl + so  
lution? + + + 0.

## Calculations for Solutions Worksheet and Key

Solutions Worksheet.

1) Why does water have such a low vapor pressure? Explain. The hydrogen bonds in

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Worksheet 1

water are strong enough that they keep molecules from leaving the surface of the liquid and entering the vapor phase. 2) Give one example of surface tension you're familiar with, and one example of a surfactant around your house.

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answers that we will  
completely offer.

## **Molarity Worksheet Answers -**

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WORKSHEET: SOLUTION  
S AND COLLIGATIVE  
PROPERTIES SET A: 1.

Find the molarity of all  
ions in a solution that  
contains 0.165 moles  
of aluminum chloride in  
820. ml solution.

Answer:  $[Al^{3+}] = 0.201$   
M ,  $[Cl^-] = 0.603$ M. 2.

Find the molarity of



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each ion present after mixing 27 ml of 0.25 M  $\text{HNO}_3$  with 36 ml of 0.42 M  $\text{Ca}(\text{NO}_3)_2$   
(Note: There is no ...

### **Worksheet\_Colligative.pdf - WORKSHEET: SOLUTIONS AND ...**

214.2g  $\text{OsF}_3 \times 1 \text{ mol}$   
 $\text{OsF}_3 = 12.9 \text{ M OsF}_3$ .  
0.0673 L soln 247.23 g  
 $\text{OsF}_3$ . Calculate the molarity if a flask contains 1.54 moles potassium sulfate in 125 ml of solution. 1.54

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mol  $K_2SO_4 = 12.3 M$

$K_2SO_4$ ... Molarity Answers

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