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## **IB Study Notes - Free IB Notes**

After that, I figured you guys might really need a practical guide on this. Yes, there is stuff on the internet that quotes the criteria and kind-of superficially tells you what they want to see, but nobody's really telling you HOW to do it. Anyways, hope this helps! This guide is most definitely applicable to IB



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Biology and Chemistry.

## **The Science IA Guide the IB Probably Doesn't Want You to ...**

Atom. Atoms consist of a three subatomic particles Proton; Neutron; Electron; Atomic Number  $Z$ : number of protons in the nucleus of an atom of an element. Mass Number  $A$ : number of

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protons + number of neutrons Relative atomic mass  $A_r$ : Ratio of the average mass of the atom to the unified atomic mass unit; Atomic Mass Units (AMU): 1/12th of the mass of a carbon - 12 atom in its ground state.

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